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Solution Chemistry Pogil

To obtain pH, we could calculate $[H_3O^+] = K_w / [OH^-]$, and then get pH; or we could calculate pOH from our value of $[OH^-]$, and get pH from $pH = 14.000 - pOH$. Taking the second approach, we have $pOH = -\log(1.48 \times 10^{-3}) = 2.830$ $pH = 14.000 - 2.830 = 11.170$.

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The pH of the solution will be governed by the following base hydrolysis equilibrium, which is just the reverse of the neutralization reaction: $A^- + H_2O \rightleftharpoons HA + OH^-$ K_b for A^- , which may be calculated from K_a of HA is $K_b = K_w / K_a = 1.00 \times 10^{-14} / 1.00 \times 10^{-5} = 1.00 \times 10^{-9}$ Making the usual assumptions we can calculate $[OH^-]$ and pOH to obtain pH:

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Calculating pH and pOH - High School Chemistry

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Chem 116 POGIL Worksheet - Week 9 - Solutions Weak Acid

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Calculating Ph Packet Answers Pogil Answer

Now we can find the pOH. The sum of the pH and the pOH is always 14. The pOH of the solution is 7.8. Alternatively, a shortcut can be used to estimate the pH. If is in the form , then pH is roughly . For this question, this shortcut gets us a pH of 6.4, which produces a pOH of 7.6; very close to the real answer!

Pogil Answer Key For Calculating Ph

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